

2.5 C4 Poly Part 2

December 6, 2023 10:20 AM

Name: _____

Science 9- Chemistry Topic 2.5 - Concept 4b: Polyatomic ions are made up of more than one atom. (Read pages 164-165)

Naming Compounds using the "Reverse Criss-Cross" Method

Steps	$K_2(SO_4)_1$	$Mn(NO_3)_2$
1. "Un-criss-cross" the subscripts back into the superscripts with charges.	$K^{1+} SO_4^{2-}$ ✓	$Mn^{2+} NO_3^{1-}$
2. Then look at the NON-METAL ION CHARGE after the "un-criss-cross", if it is NOT CORRECT , then multiply the superscripts by the proper number so the non-metal ion ends up with the right charge. The resulting charge on the metal ion is now correct!	only 1 version K^+ check back of table	check → Exist ✓
3. Write the name of the compound.	Potassium Sulfate	Manganese (II) Nitrate

Formula	Name of Compound	Formula	Name of Compound
$AlPO_4$		$Pb(OH)_4$	
KNO_3		$Cu(ClO_3)_2$	
$NaHCO_3$		$FeSO_4$	
$CaCO_3$ $Ca^{2+} CO_3^{2-}$ $Ca^+ CO_3^-$ x Calcium Carbonate		$SnSO_4$	
$Mg(OH)_2$		$Pb(NO_3)_2$	
Na_2CrO_4		$CuHSO_3$ $Cu^+ HSO_3^-$ ✓ Copper (I) Bisulfite	
$Ba(CN)_2$		$MnSO_3$	
NH_4NO_3		$Au_2(Cr_2O_7)_3$	
Na_2SO_3		$Fe(CH_3COO)_3$	
$Al(MnO_4)_3$		$HgHCO_3$	

DATE:

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TOPIC 2.5**Writing Names and Formulas of Compounds with a Polyatomic Ion****BLM 2.5-9**

Name of metal ion	Name of non-metal ion	Formula of compound	Name of compound
copper(II)	hydroxide	$\text{Cu}(\text{OH})_2$	copper(II) hydroxide
			aluminum nitrite
			lead(IV) dichromate
gold(I)	chromate		
strontium	chlorate		
iron(III)	sulfate		
copper(I)	carbonate		
		AgClO_2	
rubidium	phosphate		
		NaHCO_3	
calcium	phosphite		
			magnesium acetate
			manganese(IV) sulfite
		LiMnO_4	